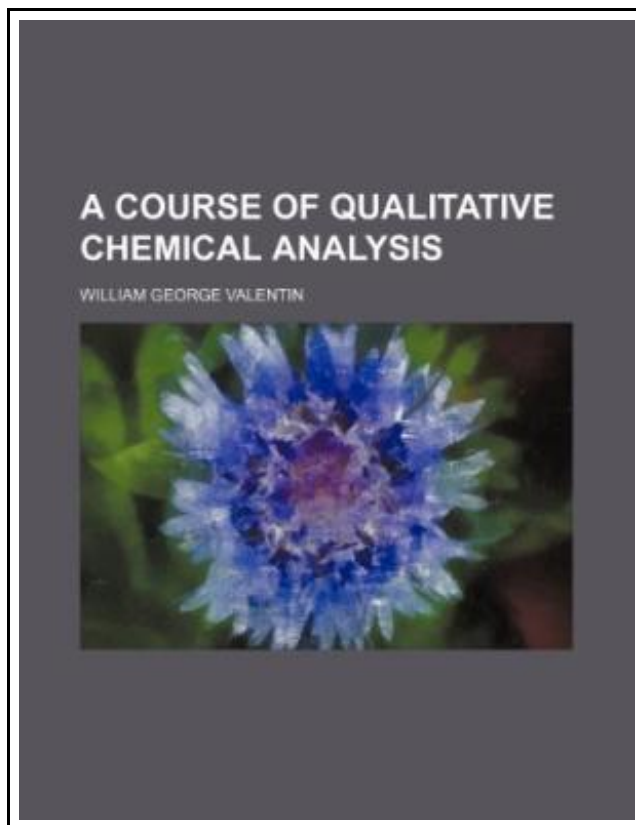


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Here is the best publication i have got go through until now. It is actually writter in simple phrases and never hard to understand. I realized this publication from my dad and i suggested this ebook to find out.

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A COURSE OF QUALITATIVE CHEMICAL ANALYSIS



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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1876 Excerpt: .of chromic acid, and indirectly of hydrochloric acid, is inferred from the yellow precipitate which a plumbic salt produces with their solutions. This reaction is employed for recognizing a chloride in the presence of a bromide. The latter yields bromine when treated similarly, and a colourless solution of NH_4Br . The presence of an iodide, yielding free iodine, would give rise to the formation of explosive black insoluble iodide of nitrogen, on the addition of NH_3 ; the reaction can, therefore, only be performed after the previous-removal of the iodide. REACTiOnS In tHE WEt WAY. A Solution Of Sodid Chloride is employed. All chlorides may be divided into: --1st. Chlorides insoluble in water, such as AgCl and Hg_2Cl_2 . which are quite insoluble; PbCl_2 , soluble in 135 parts of cold water, readily soluble in boiling water; and Cu_2Cl_2 , almost insoluble in water and in dilute sulphuric acid; soluble in ammonia, HCl and NaCl solutions. AuCl and PtCl_2 , are almost insoluble in water. 2nd. Chlorides which are soluble in water, comprising all other chlorides. The majority of the chlorides of the most powerfully basifous metals, as K , Na , Ba , etc., are less soluble in aqueous HCl , than in OH_2 . The chlorides of the less basifous metals (with the exception of Hg and Ag) are decomposed by water, with formation of free acid and metallic oxides, ex. gr. AsO_3 , or of some insoluble oxychloride (SbO_3 , BiCl_3), soluble, however, in excess of HCl . In order to ascertain the presence of a...



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